

The **Periodic Table** displays the names and symbols of all the **elements** we have discovered which are organised by their **chemical properties** and their **physical properties**. There are about 100 naturally occurring elements.

Physical properties

The **physical properties** of an element describe how a substance behaves generally.
(E.g., **conductor** of electricity, **dense**, **conductor of heat**, **shiny**, **malleable**, **sonorous**, **high melting and boiling points**)

Chemical properties

The **chemical properties** of an element describe how a substance behaves in chemical reactions.
For example, **how reactive it is**, **what other substances it reacts with**, and **the products it forms in reactions**.

Metals

- good conductors of heat and electricity
- shiny
- malleable and ductile
- sonorous when solid
- most have high melting and boiling points
- some metals react with oxygen to form metal oxides

Group 1

- called the **alkali metals**
- like all other metals but are very **reactive**
- react vigorously (strongly) with water
- get more reactive as you go down the group
- lower melting points than most other metals
- melting points decrease down the group
- always produce a metal hydroxide and hydrogen gas when reacted with water

1	2																	0/8
Li lithium	Be beryllium	H hydrogen															He helium	
Na sodium	Mg magnesium															Ne neon		
K potassium	Ca calcium	Sc scandium	Ti titanium	V vanadium	Cr chromium	Mn manganese	Fe iron	Co cobalt	Ni nickel	Cu copper	Zn zinc	Ga gallium	Ge germanium	As arsenic	Se selenium	Br bromine	Kr krypton	
Rb rubidium	Sr strontium	Y yttrium	Zr zirconium	Nb niobium	Mo molybdenum	Tc technetium	Ru ruthenium	Rh rhodium	Pd palladium	Ag silver	Cd cadmium	In indium	Sn tin	Sb antimony	Te tellurium	I iodine	Xe xenon	
Cs caesium	Ba barium	La lanthanum	Hf hafnium	Ta tantalum	W tungsten	Re rhenium	Os osmium	Ir iridium	Pt platinum	Au gold	Hg mercury	Tl thallium	Pb lead	Bi bismuth	Po polonium	At astatine	Rn radon	
Fr francium	Ra radium																	

■ solids ■ liquids ■ gases at room temperature

This version of the Periodic Table does not include every discovered element.

Group 7

- called the **halogens**
- generally very reactive
- generally the opposite of Group 1
- melting point increases down the group while reactivity decreases.
- take part in **displacement reactions**, where an element from higher up the group takes the place of one from lower down the group in a compound.

For example: $\text{potassium iodide} + \text{chlorine} \rightarrow \text{potassium chloride} + \text{iodine}$

- columns are called **groups**
 - rows are called **periods**
- Elements in a group normally have similar properties, meaning chemists can predict properties of elements based on their group.

Non-metals

- often have properties the opposite of metals
- low boiling points, so are often gases at room temperature
- poor conductors of electricity and heat
- dull in appearance
- low density
- **brittle** and not sonorous
- some metals react with oxygen to form metal oxides

Group 0

- called the **noble gases**
- very unreactive
- low boiling points, so are gases at room temperature
- like the halogens, their boiling points increase down the group

Key words

Make sure you can write definitions for these key terms.

alkali metal brittle conductor chemical property dense displacement reaction element group halogen malleable metal noble gas non-metal
period Periodic Table physical property sonorous reactive

